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journal of human kinetics volume 42/2014, 73-79 doi: 10.2478/hukin-2014-0062 73 section i - kinesiology 1 - department of biomechanics, university school physical education, wroclaw, poland. authors submitted their contribution of the article to the editorial board. **chemical kinetics reaction rates - csus** - 10 recognizing a first order process: $a \rightarrow \text{products}$ whenever the conc. of a reactant falls off exponentially, the kinetics follow first order. $k t \ln [a]_0 - k t \ln [a]_t = -k t$ a plot of $\ln [a]$ versus time (t) is a straight line with slope -k and intercept $\ln [a]_0$ $\ln [a]_t = \ln [a]_0 - k t$ = - determining the rate constant for a first order process

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